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Class 12 Important Questions for Chemistry - Electrochemistry

Two half-reactions of an electrochemical cell are given below : $\text{MnO}_4^- (\text{aq}) + 8\text{H}^+ (\text{aq}) + 5\text{e}^- \rightarrow \text{Mn}^{2+} (\text{aq}) + 4\text{H}_2\text{O} (\text{l})$, $E^\circ = 1.51 \text{ V}$. $\text{Sn}^{2+} (\text{aq}) \rightarrow \text{Sn}^{4+} (\text{aq}) + 2\text{e}^-$, $E^\circ = + 0.15 \text{ V}$. Construct the redox equation from the standard potential of the cell and predict if the reaction is reactant favoured or product favoured.

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