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Fig. 17.2, page 506 - on the left, the system (the people) gain heat from it's surroundings (the fire) this is . endothermic (q is positive) On the right, the system (the body) cools as perspiration evaporates, and heat flows to the surroundings. this is . exothermic (q is negative)

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state. chemical potential energy. energy stored in chemical bonds of a substance.

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